and from about 225 to 290 minutes, which naturally must have caused longer course of the "heat" and shrinkage of charge. . One further sees that the initial temperatures were low, in that quite as in example 8996 the carbon in the first 20 minutes (corresponding to 85) only consumed very little oxygen and left over the lion's share to the remaining reducing agents.
In this " heat " there were rendered available
0-0793 X 20,808 v r , .. f. . . ------- _i—ziz5'5
kilogrammes 0 per minute,
>J«L/U
, .T , , , , 20,808x1-0487 ni> , ., and the steel proiliiced=
—.....-t>())-----.....=73 kilogrammes per
minute, thus so much less compared with the previous example that the smaller charge weight alone is not sufficient explanation of the difference.
The following example (Table X) shows the course of a charge in which one sought to make the ore additions in small amounts corresponding to the requirements of the intervals of time. The chemical course is given in the table.
Here the slag during the entire course of the process is quite low in iron, excepting perhaps the periods from llw beginning up to 40 minutes and between 105 and 175 minutes.
Per <vnL
With the ore ('23'9 per cent.) ihoro was introduced into
the furnace "JKc 2,T!) X(H)8X °*? • - - -PC contained in final slag 18-
3x0vl007 . . . :
Thus iron reduced . . . . . . :.
There were removed reducing agents :—
3-55 per cent. C+l'2-I per cent. Bi+2'41 per cent. Mn~f
0*13 per cent. P.....: :
Thus giving rise to a charge growth of .
The ore charge, of 231) per cent, for the period of time of 220 minutes, again analogously to the first example, Charge 8,996, had been greater than the theoretical amount which would have been only 21/4 per cent. On this ground the charge growth of I'll per cent, is greater than that calculated theoretically on the assumed formula, which only amounts to 6 per cent. Unfortunately, the composition of the linal